GASES

P ⇒ collisions (#, force)

T ⇒ $\bar{x}$ velocity

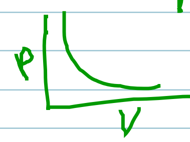
Same gas → ΔT

gas mixture ΔMm

$PV = nRT$ → $P = \frac{n}{V}RT$ ⇒ $P = MRT$ ($P \propto M$)

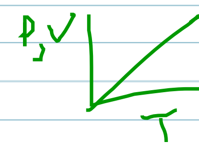
↓ $d = \frac{m}{V}$ ⇒ $d = \frac{P \cdot M_m}{RT}$ → $M_m = \frac{dRT}{P}$ ($P \propto d$)

$\frac{PV}{T} = nR$ $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$ $P_1V_1 = P_2V_2$



$\frac{V_1}{T_1} = \frac{V_2}{T_2}$

$\frac{P_1}{T_1} = \frac{P_2}{T_2}$



* collecting over H_2O

$P_T = P_{H_2O} + P_{gas}$

Dalton's $P_T = P_1 + P_2 + P_3 + \dots$

$X_A = \frac{n_A}{n_T}$ $P_A = X_A P_T$ $P \propto n$

V is same   * coefficients *

in a gas   interparticle distance ↑↑↑   IMF's ≈ 0

P↓ if IMF's ↑   V↓ if molecules are large

---

measure   mass

P, V, T } n $M_m = \frac{mass}{n}$

GASES

P ⇒ collisions (#, force)

T ⇒ $\bar{x}$ velocity

same gas → ΔT

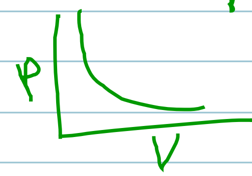
gas mixture ΔMm

$PV = nRT$ → $P = \frac{n}{V}RT$ ⇒ $P = MRT$ ($P \propto M$)

→ $d = \frac{m}{V}$ ⇒ $d = \frac{P \cdot M_m}{RT}$ → $M_m = \frac{dRT}{P}$ ($P \propto d$)

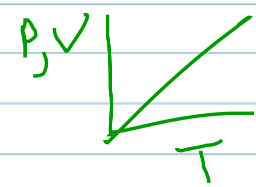
$\frac{PV}{T} = nR$ $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$

$P_1V_1 = P_2V_2$



$\frac{V_1}{T_1} = \frac{V_2}{T_2}$

$\frac{P_1}{T_1} = \frac{P_2}{T_2}$



* collecting over H_2O

$P_T = P_{H_2O} + P_{gas}$

Dalton's $P_T = P_1 + P_2 + P_3 + \ldots$

$X_A = \frac{n_A}{n_T}$ $P_A = X_A P_T$ $P \propto n$

V is same   * coefficients *

in a gas   interparticle distance ↑↑ ↑   IMF's ≈ 0

P↓ if IMF's ↑   V↓ if molecules are large

---

measure   mass

P, V, T ⟩ n $M_m = \frac{mass}{n}$