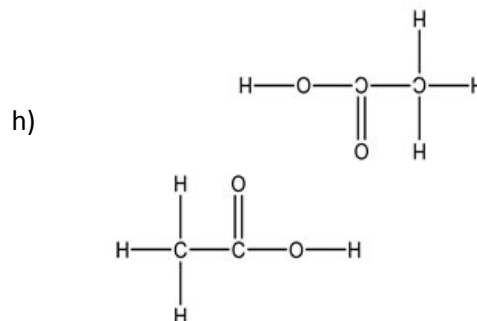
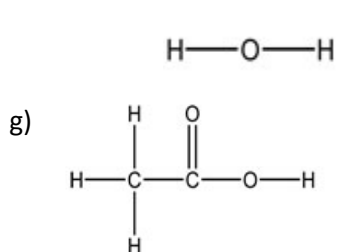
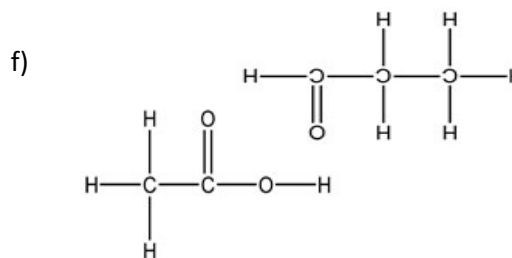
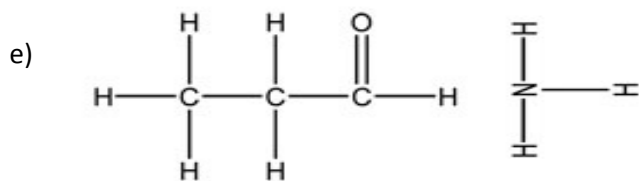
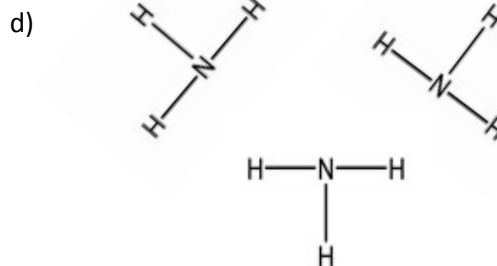
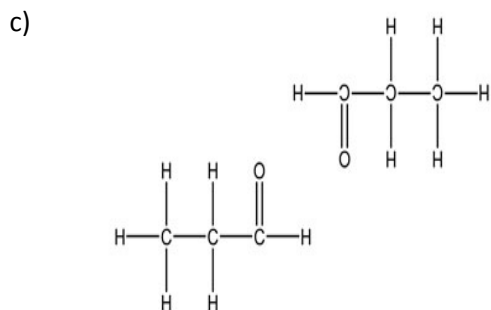
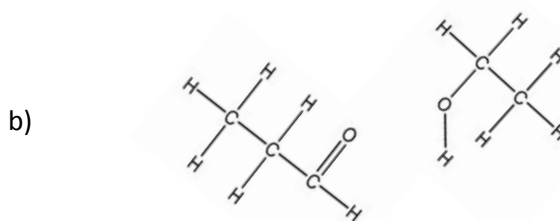
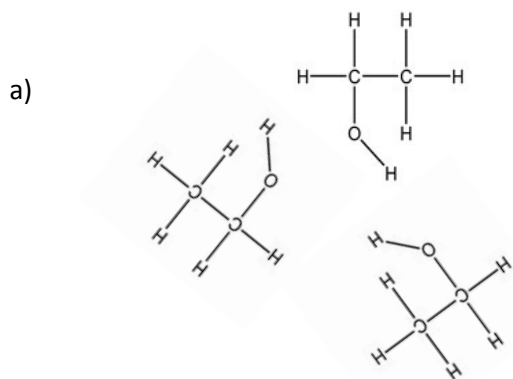
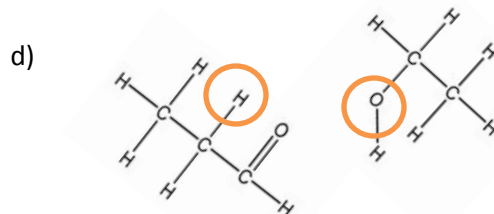
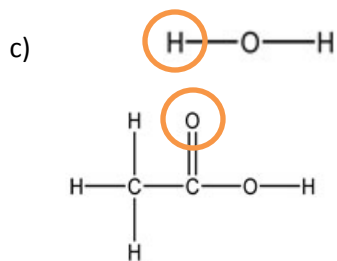
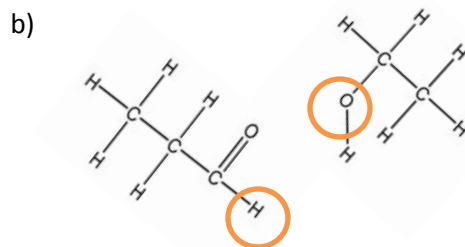
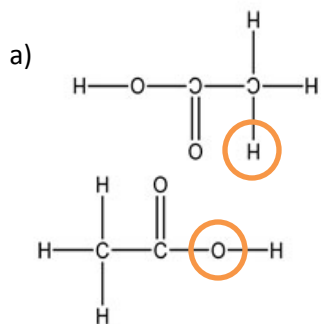


Intermolecular Forces practice

- 1) Draw a correct representation of hydrogen bonding between the following molecules. Write "no hydrogen bonding" if hydrogen bonding cannot occur between the molecules.



2) Can the circled two atoms hydrogen bond? Why or why not?



3) True or false, if false, rewrite the sentence so it is true.

a) Hydrogen bonding can occur within a molecule that has hydrogen bonded to nitrogen, oxygen, or fluorine.

b) Hydrogen bonding can occur between the nucleus of an oxygen atom and a lone pair of a hydrogen atom.

c) Hydrogen bonding is a stronger force than an ionic bond.

4) Draw the interactions between liquid water molecules and dissolved sodium chloride. (Will the positive or negative ends of water line up with the sodium ion – will the positive or negative ends of water line up with the chloride ion?)

5) Draw the interactions between $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$ and ammonia (NH_3). Recall that all carbons have 4 total bonds.