

Nomenclature Practice: Binary Ionic Compounds

name: _____

Name the following compounds:

1. NaBr _____
2. BaS _____
3. MgCl₂ _____
4. CaS _____
5. Ag₂S _____
6. AlF₃ _____
7. Cs₂O _____
8. AlP _____
9. Ca₃P₂ _____
10. KF _____

11. ZnI₂ _____
12. NaBr _____
13. AgF _____
14. BeI₂ _____
15. Sr₃N₂ _____
16. Al₂S₃ _____
17. MgO _____
18. Na₂O _____
19. RbCl _____
20. Ba₃P₂ _____

Write the formulas for the following compounds:

21. sodium bromide _____
22. barium sulfide _____
23. magnesium chloride _____
24. calcium sulfide _____
25. silver sulfide _____
26. lithium fluoride _____
27. cesium oxide _____
28. aluminum phosphide _____
29. rubidium sulfide _____
30. potassium fluoride _____

31. zinc iodide _____
32. sodium bromide _____
33. aluminum oxide _____
34. beryllium iodide _____
35. strontium nitride _____
36. aluminum sulfide _____
37. magnesium oxide _____
38. silver oxide _____
39. rubidium chloride _____
40. barium phosphide _____

Write out the Lewis dot structure *and* determine the formula for the ionic compound made by the following combinations of atoms, using arrows and dots to show where the electrons come from and go to:

a) Na & I

Formula = _____

b) Al & S

Formula = _____

c) Sr & Cl

Formula = _____

d) Be & N

Formula = _____